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Applied Chemistry Volume 1; Water, Detergents, Textiles, Fuels, Etc

By Charles Kenneth Tinkler

Not Avail, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1920 Excerpt: .concentrated solutions, but tends to decompose in the presence of small quantities of impurities, such as are usually found in commercial preparations. Determination of the Amount of Hydrogen Peroxide or of Available Oxygen in a Solution of Hydrogen Peroxide. The amount of hydrogen peroxide, or of available oxygen, in a solution of hydrogen peroxide, can be determined by several different methods: --I. By Determining the Amount of Iodine Liberated from an Acidified Solution of Potassium Iodide.-- Hydrogen peroxide, like other oxidising agents, liberates iodine from an acidified solution of potassium iodide, and if considerable excess of sulphuric acid is present, the reaction may be made use of quantitatively. $H_2O_2 + 2HI \rightarrow 2H_2O + I_2$ The iodine thus liberated can be titrated with decinormal sodium thiosulphate solution, as described under bleaching powder (seep. 137). The above equation shows that 17 parts H_2O_2 ...


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